

Scipad year 11 answers selected elements Updated Version

scipad year 11 answers selected elements is a crucial resource for students studying science in Year 11. It provides comprehensive answers and insights into key elements and topics covered in the curriculum, enabling students to deepen their understanding and excel in their assessments. Whether you're preparing for exams, completing coursework, or seeking clarification on complex concepts, having access to well-structured, accurate answers related to selected elements in Scipad can significantly enhance your learning experience.

Understanding the Importance of Selected Elements in Scipad Year 11

What is Scipad?

Scipad is an educational platform designed to assist Year 11 students in mastering science topics across various disciplines, including Chemistry, Physics, and Biology. It offers practice questions, detailed answers, and interactive resources aligned with the Australian Curriculum.

Why Focus on Selected Elements?

In science, understanding the properties and behaviors of elements is fundamental. Selected elements often serve as foundational concepts that underpin broader topics such as chemical reactions, periodic table trends, and biological processes. Mastering these elements enables students to:

- Build a strong conceptual framework
- Succeed in assessments and exams
- Connect theoretical knowledge to real-world applications

Key Elements Covered in Scipad Year 11 Answers

- Hydrogen (H)

Properties and Significance

Hydrogen is the simplest and most abundant element in the universe. It plays a vital role in various

chemical reactions, including combustion and the formation of water.

Common Applications

- Fuel cells
- Industrial synthesis (e.g., ammonia production)
- Rocket fuel

Student Focus Areas

- Electron configuration of hydrogen
- Hydrogen bonding and molecular structure
- Reactions involving hydrogen

- Carbon (C)

Properties and Significance

Carbon is renowned for its versatility, forming the basis of organic chemistry. Its ability to form four covalent bonds makes it integral to countless compounds.

Key Compounds

- Hydrocarbons (alkanes, alkenes, alkynes)
- Carbohydrates, lipids, proteins (biological molecules)

Student Focus Areas

- The carbon cycle
- Structural formulas of organic molecules
- Isomerism in carbon compounds

- Oxygen (O)

Properties and Significance

Oxygen is essential for respiration and combustion. It also plays a critical role in oxidation-reduction reactions.

Common Reactions

- Combustion of hydrocarbons
- Formation of oxides
- Photosynthesis and respiration processes

Student Focus Areas

- Electron transfer in oxidation/reduction
- The role of oxygen in biological systems
- Properties of dioxygen (O₂) and ozone (O₃)

- Nitrogen (N)

Properties and Significance

Nitrogen makes up about 78% of Earth's atmosphere. It is vital for amino acids and nucleic acids, the building blocks of life.

Key Applications

- Production of ammonia for fertilizers
- Nitrogen fixation processes
- Inert atmosphere applications

Student Focus Areas

- The nitrogen cycle
- The Haber process
- Biological nitrogen fixation

- Sodium (Na)

Properties and Significance

A highly reactive metal, sodium is essential in biological systems and industrial processes.

Common Reactions

- Reactivity with water producing sodium hydroxide and hydrogen
- Formation of sodium compounds (e.g., NaCl)

Student Focus Areas

- Electron configuration of sodium
- Alkali metal properties
- Sodium's role in the human body

How to Use Scipad Year 11 Answers for Selected Elements Effectively

Step 1: Review Basic Concepts

Start by understanding the fundamental properties of each element, including atomic structure, reactivity, and common compounds.

Step 2: Study Detailed Answer Explanations

Leverage the comprehensive answers provided to clarify complex topics, such as chemical bonding, periodic trends, and biological significance.

Step 3: Practice with Related Questions

Use practice questions linked to each element to test your knowledge. Attempt to answer without looking at the solutions first, then review your responses using the detailed answers.

Step 4: Focus on Application and Real-World Connections

Relate the elements to real-world applications and environmental processes to deepen understanding and retention.

Step 5: Regular Revision

Consistently revisit the answers and summaries to reinforce learning, especially before exams or assessments.

Key Strategies for Maximizing Learning from Scipad Year 11 Answers

- Active Engagement

- Take notes while studying answers
- Highlight key points and concepts
- Summarize answers in your own words

- Concept Mapping

Create visual diagrams linking elements to related topics, reactions, and applications. This helps in understanding connections and relationships.

- Group Study and Discussion

Discuss answers and concepts with peers to gain different perspectives and clarify doubts.

- Seek Clarification

If any answer or concept remains unclear, consult teachers or additional resources to ensure a thorough understanding.

Additional Resources to Complement Scipad Year 11 Answers

- Textbooks: Refer to your prescribed science textbooks for detailed explanations
- Online Educational Platforms: Use videos and tutorials for visual learning
- Practice Tests: Regularly test yourself with past papers and quizzes
- Laboratory Experiments: Engage in practical activities to reinforce theoretical knowledge

Conclusion

scipad year 11 answers selected elements serve as an invaluable resource for students aiming to

excel in their science studies. By focusing on key elements such as hydrogen, carbon, oxygen, nitrogen, and sodium, students can build a solid foundation in understanding chemical properties, reactions, and real-world applications. Effective utilization of detailed answers, combined with active learning strategies, can significantly improve comprehension, retention, and exam performance. Remember, mastering these elements not only aids academic success but also enhances your appreciation of the fundamental building blocks of the natural world. Embrace the resources available through Scipad and approach your studies with curiosity and confidence for a successful Year 11 science journey.

Scipad Year 11 Answers Selected Elements: A Comprehensive Guide for Students and Educators

Navigating the complexities of Scipad Year 11 answers selected elements can often feel like deciphering a code designed for seasoned experts. For students aiming to excel in their assessments and educators seeking to understand the nuances of student responses, a detailed breakdown of these answers is invaluable. In this guide, we will explore the key elements involved in Scipad Year 11 answers selected elements, analyze common patterns, and provide strategies to interpret and improve responses effectively.

Understanding the Role of Selected Elements in Scipad Year 11 Answers

Selected elements refer to specific parts of a question or task that students are expected to focus on when formulating their answers. These elements serve as the critical points that demonstrate understanding, analytical ability, and application skills. Recognizing and addressing these elements accurately is essential in achieving high marks.

What Are Selected Elements?

- Definition: The specific aspects or components within a question that students are instructed to consider or incorporate in their responses.
- Purpose: To guide students in focusing their answers, ensuring all facets of the question are addressed.
- Examples: In a science question, selected elements might include variables, scientific principles, or experimental procedures.

The Significance of Selected Elements in Year 11 Responses

Accurately identifying and responding to selected elements is central to demonstrating comprehension and critical thinking. It enables students to:

- Showcase understanding of core concepts.
- Address all aspects of multi-part questions.
- Structure responses logically around key points.
- Meet assessment criteria related to depth and breadth of knowledge.

For educators, analyzing how students approach these selected elements offers insight into their conceptual grasp and potential areas for instructional improvement.

Common Patterns in Selected Elements for Year 11 Answers

When analyzing student responses, certain patterns tend to emerge regarding how selected elements are addressed:

- Explicit Identification and Focus

Many high-scoring responses explicitly mention the key elements early in their answer, demonstrating clarity of understanding.

- Example: "The key element here is the effect of temperature on enzyme activity..."

- Integration of Elements with Explanation

Students often integrate these selected elements seamlessly with explanations, showing depth of insight.

- Example: Explaining how increased temperature can cause enzyme denaturation, affecting the rate of reaction.

- Linking Elements to Broader Concepts

Effective answers connect selected elements to broader principles or real-world applications, illustrating comprehensive understanding.

- Example: Relating the effect of pH on enzyme structure to its activity in biological systems.

Strategies for Interpreting Selected Elements in Student Answers

For educators and students alike, effective interpretation of selected elements involves careful analysis of responses.

For Educators:

- Identify the Focus: Determine if the student has addressed all the key elements specified in the question.
- Assess Depth of Understanding: Look for explanations that go beyond surface-level descriptions.
- Check for Integration: Evaluate whether the student connects the selected elements logically and coherently.
- Note Omissions: Recognize any missing elements that could indicate gaps in understanding.

For Students:

- Highlight Key Elements: Before answering, identify the specific elements the question emphasizes.
- Structure Responses: Use the selected elements as headings or focal points within your answer.
- Provide Explanations: For each element, include detailed explanations, examples, or applications.
- Review Against Question: Ensure all selected elements are addressed thoroughly before submission.

Common Challenges and How to Overcome Them

Despite the importance of selected elements, students often face challenges in addressing them effectively.

Challenge 1: Misidentifying Elements

- Solution: Carefully read the question and underline or note all key terms and instructions.

Challenge 2: Superficial Responses

- Solution: Develop detailed explanations for each selected element, including reasons, effects, or examples.

Challenge 3: Omitting Critical Elements

- Solution: Create a checklist of all elements before answering and cross-reference your response for completeness.

Enhancing Responses Through Effective Use of Selected Elements

To maximize marks and demonstrate mastery, students should:

- Explicitly state the element in their answer to clarify their focus.
- Use precise terminology relevant to the subject area.
- Support statements with examples, diagrams, or data where applicable.
- Connect elements to show relationships and broader concepts.

Practical Tips for Students:

- Practice analyzing past questions to identify common selected elements.
- Develop templates or frameworks that organize responses around key elements.
- Seek feedback on how well responses address the selected elements and adjust accordingly.

The Role of Teachers in Guiding Students on Selected Elements

Teachers play a crucial role in helping students master the art of addressing selected elements:

- Explicit Instruction: Teach students how to identify and interpret key elements within questions.
- Model Responses: Demonstrate high-quality answers that effectively incorporate selected elements.
- Provide Feedback: Highlight how well students addressed the elements and suggest improvements.
- Design Practice Questions: Create exercises that focus on recognizing and responding to selected elements.

Conclusion

Mastering Scipad Year 11 answers selected elements is fundamental to academic success at the Year 11 level. Whether you are a student aiming to improve your response quality or an educator seeking to enhance assessment strategies, understanding how to identify, interpret, and address these elements is key. By focusing on clarity, depth, and integration, responses can truly demonstrate the learner's grasp of essential concepts, leading to higher achievement and a more thorough understanding of the subject matter.

Remember, the goal is not just to answer but to answer thoughtfully, highlighting the critical elements and weaving them into a comprehensive, insightful response.

Question What are the key elements to focus on when studying the SciPad Year 11 answers for selected elements? The key elements include understanding atomic structure, properties of specific elements, their placement in the periodic table, and their common compounds, as well as practicing application-based questions to solidify understanding. **How can I effectively use SciPad Year 11 answers to improve my understanding of selected elements?** Use the answers to review correct responses, understand the reasoning behind each answer, and compare your responses to identify areas for improvement. Practice consistently with these solutions to build confidence. **Which selected elements are most commonly covered in SciPad Year 11 answers?** Commonly covered elements include Hydrogen, Carbon, Oxygen, Nitrogen, Sodium, and Chlorine, as they are fundamental to understanding basic chemistry concepts and periodic table trends. **Are the SciPad Year 11 answers for selected elements suitable for exam preparation?** Yes, they are designed to align with the curriculum and can serve as valuable resources for revision and practice, helping students prepare effectively for their exams. **How can I identify which answers to focus on in the SciPad Year 11 solutions for selected elements?** Focus on questions you find challenging, review the detailed solutions provided, and cross-reference with your textbook or notes to ensure a comprehensive understanding of each element's concepts. **What strategies can I use to memorize key facts about selected elements using SciPad Year 11 answers?** Use active recall by testing yourself on the answers, create flashcards for important facts, and regularly review the solutions to reinforce your memory and understanding of each element.

Related keywords: scipad, year 11, answers, selected elements, chemistry, periodic table, data analysis, chemical properties, educational resources, student guide

Elements And Compounds Comprehension

Elements and compounds comprehension is a fundamental aspect of understanding the building blocks of matter in the field of chemistry. Grasping the differences between elements and compounds, their properties, and how they interact is essential for students, educators, and anyone

interested in the sciences. This article aims to provide a comprehensive overview of these core concepts, exploring their definitions, characteristics, examples, and significance in the natural world.

Understanding Elements

What Are Elements?

Elements are pure substances consisting of only one type of atom. They are the simplest form of matter that cannot be broken down into simpler substances by chemical means. Each element is distinguished by its unique number of protons in the nucleus, known as the atomic number.

Characteristics of Elements

- Pure Substances: Elements are pure, meaning they contain only one type of atom.
- Unique Properties: Each element has distinct physical and chemical properties, such as melting point, boiling point, reactivity, and color.
- Atoms and Molecules: Elements can exist as individual atoms (e.g., Helium, He) or as molecules composed of two or more atoms bonded together (e.g., Oxygen, O₂).
- Periodic Table: Elements are organized systematically in the periodic table, which groups elements based on similar properties.

Examples of Elements

- Hydrogen (H)
- Carbon (C)
- Oxygen (O)
- Iron (Fe)
- Gold (Au)
- Nitrogen (N)

Understanding Compounds

What Are Compounds?

Compounds are substances formed when two or more different elements chemically bond together in fixed proportions. Unlike mixtures, compounds have specific chemical formulas that denote the element ratios, and they possess properties different from their constituent elements.

Characteristics of Compounds

- Chemical Bonding: Elements combine through chemical bonds such as covalent or ionic bonds.
- Fixed Composition: The ratio of elements in a compound is definite and constant, represented by a chemical formula (e.g., H_2O , $NaCl$).
- New Properties: Compounds often display properties different from the individual elements that form them.
- Formation and Decomposition: Compounds can be broken down into elements through chemical reactions, such as electrolysis or heating.

Examples of Compounds

- Water (H_2O)
- Sodium chloride ($NaCl$)
- Carbon dioxide (CO_2)
- Methane (CH_4)
- Glucose ($C_6H_{12}O_6$)

Differences Between Elements and Compounds

Aspect	Elements	Compounds
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Definition	Pure substances with one type of atom	Substances formed by chemically bonded atoms of different elements
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Composition	Only one element	Two or more different elements
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Chemical Formula	Usually not written or simple (e.g., O_2)	Specific formula (e.g., H_2O)
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Properties	Unique, characteristic of the element	Different from constituent elements
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Breakdown Process	Cannot be broken down chemically	Can be broken down into elements via chemical reactions
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The Significance of Elements and Compounds

Understanding elements and compounds is crucial for several reasons:

- Foundation of Chemistry: They are the basic units of matter, forming the basis of chemical

reactions and processes.

- Material Science: Knowledge about elements and compounds aids in developing new materials, medicines, and technologies.
- Environmental Impact: Recognizing how elements and compounds behave helps in understanding pollution, resource management, and sustainability.
- Biological Importance: Many biological processes depend on specific compounds, such as proteins, carbohydrates, and nucleic acids.

Methods of Identifying and Analyzing Elements and Compounds

Laboratory Techniques

- Spectroscopy: Used to identify elements based on their emission or absorption spectra.
- Chemical Tests: Reactions that produce characteristic colors or precipitates help identify specific compounds.
- Chromatography: Separates mixtures into individual components, useful in analyzing complex compounds.

Computational Methods

- Chemical Formula Prediction: Software tools can predict possible formulas for unknown substances.
- Molecular Modeling: Visualizes the structure of compounds, aiding in understanding their properties.

Applications of Elements and Compounds in Daily Life

- Medications: Many drugs are compounds designed for specific biological functions.
- Construction Materials: Elements like iron and compounds like cement are vital in construction.
- Energy Production: Combustion of hydrocarbons (compounds) fuels engines and power plants.
- Electronics: Semiconductors made from elements like silicon are crucial in modern electronics.
- Food Industry: Nutrients are compounds vital for health, such as vitamins and minerals.

Common Misconceptions About Elements and Compounds

- Elements Are Simple: While elements are basic, they can have complex structures and behaviors.

- All Compounds Are Molecules: Some compounds, like salts, form crystalline structures and are not discrete molecules.
- Elements Cannot Be Broken Down: Elements can be broken down in nuclear reactions but not by chemical means.
- Compounds Are Always Harmful: Many compounds are essential for life and beneficial, such as water and oxygen.

Conclusion

A comprehensive understanding of elements and compounds is essential to grasp the fundamental nature of matter. Elements serve as the building blocks, each with unique properties, while compounds are the chemical combinations of these elements, forming substances with new characteristics. Recognizing their differences, formation, and significance helps in appreciating the complexity and diversity of the material world. Whether in scientific research, industrial applications, or daily life, knowledge of elements and compounds underpins many innovations and solutions to global challenges.

By mastering these concepts, students and enthusiasts can develop a solid foundation in chemistry, enabling further exploration into molecular structures, chemical reactions, and the development of new materials. Continuous learning about the elements and compounds shaping our universe is crucial for advancing science and improving our quality of life.

Elements and Compounds Comprehension: An In-Depth Exploration of Matter's Fundamental Building Blocks

Understanding the nature of matter is foundational to the fields of chemistry, physics, and material science. At the core of this understanding lies the distinction between elements and compounds—two fundamental categories that define the composition and behavior of everything from the simplest gases to complex biological molecules. This comprehensive review aims to dissect the concepts of elements and compounds, elucidate their properties, their roles in the natural world, and the scientific principles underpinning their formation and classification.

Introduction to Elements and Compounds

The universe, in its vast complexity, is constructed from a finite set of building blocks. These building blocks are primarily classified into elements—pure substances consisting of only one type of atom—and compounds, which are substances formed from two or more elements chemically bonded in fixed proportions. Recognizing and differentiating between these fundamental types of matter is critical for understanding chemical reactions, material properties, and the organization of matter at the atomic level.

Defining Elements: The Building Blocks of Matter

What Are Elements?

Elements are substances composed entirely of one type of atom, distinguished by their unique number of protons, known as the atomic number. They are the simplest form of matter that cannot be broken down into simpler substances by chemical means. Each element has a specific set of properties, including atomic weight, density, melting point, and reactivity, which are characteristic and contribute to the element's identity.

Historical Perspective and Discovery

The concept of elements has evolved considerably over centuries. Early chemists identified elements through observation of natural substances and their reactions. The periodic table, developed by Dmitri Mendeleev and later refined, organizes elements based on atomic number and chemical properties, revealing periodic trends and relationships.

Classification and Properties of Elements

Elements are broadly classified into metals, nonmetals, and metalloids:

- Metals: Good conductors of electricity, malleable, ductile, often shiny (e.g., iron, copper, aluminum).
- Nonmetals: Poor conductors, brittle in solid form, diverse in physical states (e.g., oxygen, sulfur, nitrogen).
- Metalloids: Exhibit properties intermediate between metals and nonmetals (e.g., silicon, arsenic).

Elements also differ significantly in their reactivity:

- Noble gases (helium, neon, argon) are notably inert.
- Alkali metals (lithium, sodium, potassium) are highly reactive.

Atomic Structure and Isotopes

Each element's atoms are characterized by their atomic structure:

- Protons: Define the element.

- Neutrons: Vary among isotopes, affecting atomic mass.
- Electrons: Determine chemical behavior.

Isotopes are variants of an element with different neutron counts, influencing atomic weight but not chemical properties significantly.

Understanding Compounds: The Chemical Bonding of Elements

What Are Compounds?

Compounds are substances formed when two or more elements combine chemically in fixed ratios, resulting in new substances with distinct properties. The formation of compounds involves chemical bonds—either ionic or covalent—that hold atoms together.

Types of Chemical Bonds

- Ionic Bonds: Formed through electrostatic attraction between oppositely charged ions, typically between metals and nonmetals (e.g., sodium chloride, NaCl).
- Covalent Bonds: Sharing of electron pairs between atoms, common among nonmetals (e.g., water, H₂O; carbon dioxide, CO₂).

Characteristics of Compounds

- Fixed Composition: Defined by a chemical formula indicating the ratio of elements (e.g., H₂O, C₆H₁₂O₆).
- Distinct Properties: Different from constituent elements; for example, sodium is reactive and soft, while sodium chloride is stable and edible.
- Chemical Reactivity: Compounds can undergo further reactions, breaking down into elements or forming new compounds.

Classification of Compounds

Compounds are grouped based on their bonding and composition:

- Inorganic Compounds: Usually contain metals and nonmetals; include salts, minerals, acids.
- Organic Compounds: Primarily contain carbon and hydrogen; include hydrocarbons, carbohydrates, proteins.

Methods of Identification and Analysis

Distinguishing Elements from Compounds

The key differences include:

- Pure Substance: Elements are pure, while compounds are chemically combined.
- Chemical Formula: Elements typically represented by a one- or two-letter symbol; compounds have more complex formulas.
- Separation Techniques: Elements can be isolated via physical methods like electrolysis (for ions), while compounds require chemical reactions to decompose into elements.

Analytical Techniques in Modern Chemistry

- Spectroscopy: Determines atomic and molecular composition.
- Chromatography: Separates mixtures to identify constituent elements or compounds.
- Mass Spectrometry: Measures molecular weights and structures.
- X-ray Crystallography: Reveals atomic arrangements within compounds.

Significance of Elements and Compounds in Nature and Industry

Natural Occurrence

Elements and compounds are omnipresent:

- Atmosphere: Composed mainly of nitrogen (element) and oxygen (element).
- Minerals: Contain compounds like calcite (CaCO_3) and halite (NaCl).
- Biological Systems: Consist of complex compounds such as proteins, nucleic acids, and lipids.

Industrial Applications

- Materials Manufacturing: Alloys (metal mixtures), plastics (organic compounds), pharmaceuticals.
- Energy Production: Combustion of hydrocarbons, nuclear reactions involving specific elements.
- Environmental Science: Monitoring and controlling pollutants, which are often compounds or elements.

Emerging Topics and Future Directions

Advanced Material Design

Understanding elements and compounds at the atomic level enables the creation of novel materials:

- Nanomaterials: Exploit unique properties of elements and compounds at nanoscale.
- Smart Materials: Responsive to environmental stimuli, designed from specific compounds.

Chemical Synthesis and Green Chemistry

Developing sustainable methods to synthesize compounds with minimal environmental impact remains a priority, relying heavily on knowledge of element reactivity and bonding.

Elemental and Compound Databases

Technological advancements have led to comprehensive databases that facilitate research, discovery, and application of elements and compounds in various scientific fields.

Conclusion

The comprehension of elements and compounds is a cornerstone of chemistry and material science. Recognizing that elements serve as the basic building blocks of matter, while compounds are their chemically bonded assemblages, provides a framework for understanding the complexity and diversity of the material universe. Ongoing research continues to unveil the properties, behaviors, and potential applications of these fundamental substances, driving innovation across multiple scientific and industrial domains.

A thorough grasp of these concepts not only deepens scientific knowledge but also empowers technological progress, environmental stewardship, and the development of new materials that shape our future.

Question What is the difference between an element and a compound? An element is a pure substance made of only one type of atom, while a compound consists of two or more different elements chemically combined in a fixed ratio. How can you tell if a substance is an element or a compound? If a substance can be broken down into simpler substances by physical means, it is likely an element. If it can only be broken down by chemical reactions into simpler substances, it is a compound. What are some common examples of elements and compounds? Common elements include oxygen, hydrogen, and gold. Examples of compounds include water (H_2O), carbon dioxide (CO_2), and sodium chloride ($NaCl$). Why do elements have chemical symbols, and how are they

written? Elements have chemical symbols to uniquely identify each element, usually consisting of one or two letters, with the first letter capitalized and the second letter lowercase (e.g., O for oxygen, Na for sodium). What is a chemical formula, and what does it tell us about a compound? A chemical formula shows the types and numbers of atoms in a molecule of a compound, such as H_2O indicating two hydrogen atoms and one oxygen atom. How are elements and compounds represented in the periodic table? Elements are represented by their chemical symbols in the periodic table, while compounds are not listed as individual entries but can be written using chemical formulas based on the elements involved. Can elements be broken down into simpler substances? What about compounds? Elements cannot be broken down into simpler substances by chemical means; they are already the simplest form. Compounds can be broken down into their constituent elements through chemical reactions. What is the importance of understanding elements and compounds in chemistry? Understanding elements and compounds is fundamental in chemistry because it helps explain the composition, structure, and properties of matter, enabling us to predict how substances will react and interact. How do elements combine to form compounds? Elements combine through chemical bonds, such as covalent or ionic bonds, to form compounds with new properties different from the individual elements.

Related keywords: elements, compounds, chemistry, molecules, atoms, chemical formulas, periodic table, bonding, mixtures, reactions

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